Third Exam	Name (PRINT)	Last First
Chemistry 3331 November 22, 2013	Signature	Last, First
PLEASE CIRCLE CLASS TIME 10:00 AM		

Page #	Score
1. 15 pt	
2. 24 pt	
3. 18 pt	
4. 18 pt	
5. 12 pt	
6. 13 pt	

1:00 PM

Total: _____

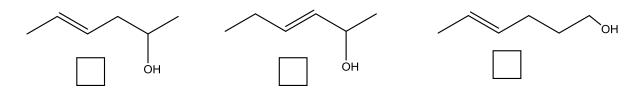
NOTE: Present your ID when you return the exam booklet

A. Nomenclature: (15 points)

Give an acceptable IUPAC name for each of the following compounds. Be sure to indicate the **stereochemistry** where appropriate.

B. Facts: Total Points = 24

1. Place the alcohols in order of increasing reactivity in an acid catalyzed dehydration. (1=least reactive, 3=most reactive) (6 points)



2. Place the compounds in order of increasing acidity. (1=weakest acid, 3=strongest acid) (6 points)

$$C \equiv C - H$$

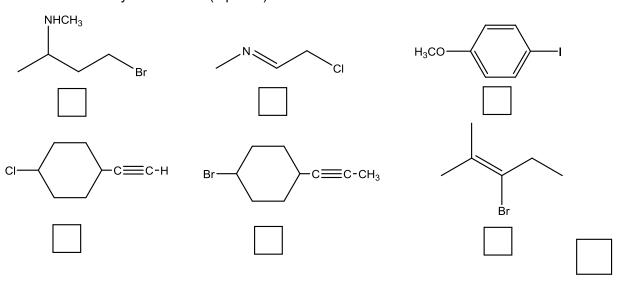
$$OH$$

$$CH_2CHCH_3$$

$$CH_2CHCH_3$$

3. Place the compounds in order of increasing acidity. (1=weakest acid, 3=strongest acid) (6 points)

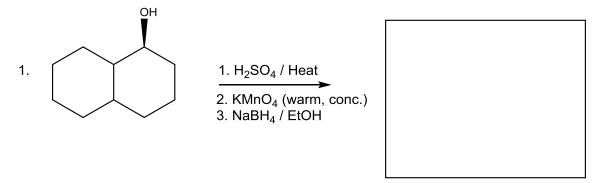
4. Place a "**Y**" in the box below any halide that will produce a useful Grignard reagent. Place an "**N**" in the box below any that will not. (6 points)

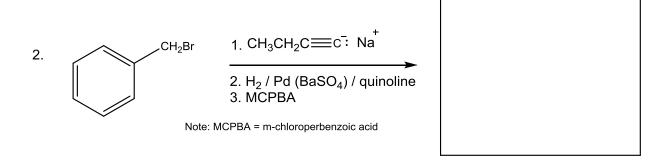


Made by Rana Shammas ranashammas93@yahoo.com

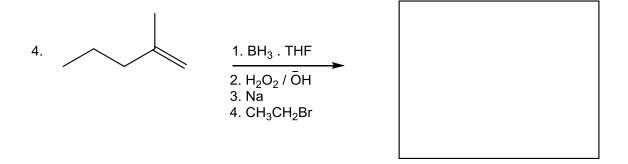
C. Reactions: Total= 36 points, 6 points each

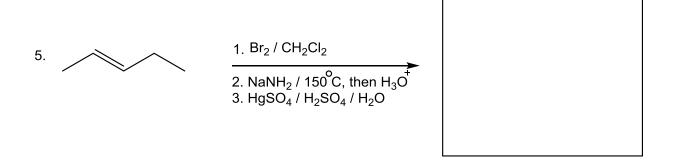
Please provide the major product or the reagents in the box. Be sure your drawing indicates stereochemistry if applicable. Partial credit is awarded only when intermediate products in a multistep reaction are shown below the reaction.

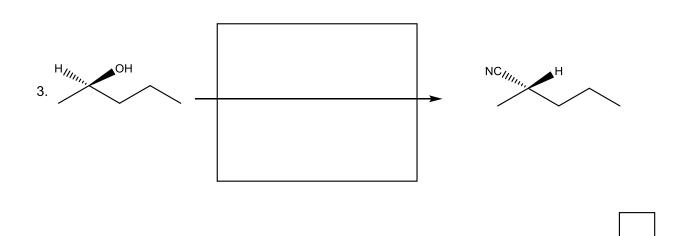




3. OH	1. HBr, ether 0 C 2. Mg / Et ₂ O	
	3. O II CH ₃ CCI, then H ₃ O	







D. Mechanisms: 12 points

The reaction below produces a mixture of products. Provide a clear mechanism to explain the formation of the product shown. Use curved arrows to indicate "electron flow". Remember to show only one step at a time. Show all intermediates and all formal charges. Do not show transition states.

E. Synthesis: 13 points

Synthesize the molecule below from alkanes, alkynes, or alcohols of **three** carbons or less, and any inorganic reagents. (Please do not include mechanisms.)

$$\begin{array}{c} \text{OH} \\ \\ | \\ \text{CH}_3\text{CH}_2\text{CH}_2 \\ - \text{C} \\ - \text{CH}_2\text{CH}_2 \\ \\ \text{CH}_2\text{CH}_3 \end{array}$$